Versatile routes to selenoether functionalised tertiary phosphines

This item was submitted to Loughborough University’s Institutional Repository by the/an author.

Citation: CUNNINGHAM, T.J. ... et al, 2010. Versatile routes to selenoether functionalised tertiary phosphines. Dalton Transactions, 39 (22), pp. 5216 - 5218.

Additional Information:

- This article was published in the journal, Dalton Transactions [© The Royal Society of Chemistry]. The definitive version is available at: http://dx.doi.org/10.1039/c0dt00004c

Metadata Record: https://dspace.lboro.ac.uk/2134/15952

Version: Accepted for publication

Publisher: © The Royal Society of Chemistry

Rights: This work is made available according to the conditions of the Creative Commons Attribution-NonCommercial-NoDerivatives 4.0 International (CC BY-NC-ND 4.0) licence. Full details of this licence are available at: https://creativecommons.org/licenses/by-nc-nd/4.0/

Please cite the published version.
Versatile Routes to Selenoether Functionalised Tertiary Phosphines

Tom J. Cunningham, Mark R. J. Elsegood, Paul F. Kelly, Martin B. Smith* and Paul M. Staniland

Received (in XXX, XXX) 1st January 2007, Accepted 1st January 2007
First published on the web 1st January 2007
DOI: 10.1039/b000000x

New selenoether functionalised tertiary phosphines, based on aryl (2a, 2b) or alkyl (4) backbones, have been synthesised and characterised. P-Se-chelation has been achieved upon complexation to square-planar PtII (3a) or PdII (3b) metal centres. For 3a and 3b, weak non covalent M···Se contacts were established using single crystal X-ray crystallography.

Tertiary phosphines continue to remain an integral tool in the design and synthesis of new metal based complexes. Hybrid tertiary phosphines, bearing additional donor atoms such as oxygen, have frequently been described as hemilabile by virtue of the soft/hard donor atom combination.1 Considerable recent interest has focused on mixed P,O- and P,S-ligands for their fascinating coordination chemistry,2 including water-soluble macrocyclic complexes using a weak-link approach,3 and catalytic applications.4 Tertiary phosphines with an additional group in the synthesis of new phosphine ligands. Herein we describe the synthesis of a new, air-stable, selenoether phosphine containing this diamondoid group, along with two further selenoether phosphines and their square-planar dichloropalladium and platinum(II) complexes.

Scheme 1

Tom J. Cunningham, Mark R. J. Elsegood, Paul F. Kelly, Martin B. Smith* and Paul M. Staniland

Versatile Routes to Selenoether Functionalised Tertiary Phosphines

The synthetic procedure to the new selenoether modified phosphines 2a and 2b was achieved using standard methodology (Scheme 1). Following a known procedure for accessing bromophosphine 1a,6 reaction of the parent cage phosphine, AdPH, with 1-bromo-2-iodobenzene in dimethylacetamide (DMA) under standard P–C coupling conditions afforded 1b in good yield (55%) as an air-stable crystalline solid.6 One noticeable difference in the preparation of 1b, opposed to 1a, was the considerably shorter reaction time to proceed to completion (typically 2 d for 1b; 5 d for 1a). In the 31P{1H} NMR spectrum of 1b a single phosphorus resonance was observed at δ(P) −29.6 ppm, some 25 ppm upfield with respect to 1a [δ(P) −4.9 ppm]. Reaction of either 1a or 1b with PhSeH and KOH, in DMA, for 5 d at ca. 160 °C afforded 2a or 2b in reasonable yields (53% and 40% respectively). 31P{1H} NMR spectroscopy of 2a and 2b showed single phosphorus resonances at δ(P) −10.0 [3J(PSe) 139 Hz] and δ(P) −36.4 ppm [3J(PSe) 201 Hz] respectively.

Crystals of 2b, suitable for X-ray crystallography, were grown by vapour diffusion of Et2O into a CDCl3 solution.7 The molecular structure (Fig. 1) of 2b revealed crystallisation of a single enantiomer and confirmed the ortho arrangement of the bulky chiral phosphadamanantyl cage and selenoether group. Metric parameters are broadly as anticipated and the transannular P(1)···Se(1) separation [3.230 Å] suggests cis-co-ordination to metal centres should be feasible. To the best of our knowledge we believe this structure determination represents the first example of a functionalised selenoether tertiary phosphine.

Fig. 1 Molecular structure of 2b. All C–H hydrogen atoms are omitted for clarity. Selected bond distances (Å) and angles (deg): C(1)···Se(1) 1.919(3), Se(1)···C(7) 1.925(3), C(8)···P(1) 1.844(3); C(1)···Se(1)···C(7) 102.4(14), Se(1)···C(7)···C(8) 119.7(2), C(13)···P(1)···C(16) 92.4(14).

To demonstrate the co-ordination potential of 2a and 2b, reaction of 1 equiv. of each ligand with either PtCl2(cod) (cod = cycloocta-1,5-diene) or PdCl2(PhCN); in CH2Cl2 at ambient temperature afforded the dichlorometal(II) complexes 3a and 3b in good yields. Crystals of 3a and 3b were grown by vapour
diffusion of diethyl ether into either a CDCl₃/(CH₃)₂SO or CH₂Cl₂ solution (Fig. 2). In both cases, P,Se-chelation is clearly evident with P(1)−M(1)−Se(1) bond angles of 90.28(3)° and 87.902(12)°. Furthermore the M(1)−Se(1) bond lengths [Pt(1)−Se(1) 2.3561(5) for 2a; Pd(1)−Se(1) 2.3656(2) for 2b] are consistent with selenoether coordination. The M(1)−Se(1)−C(7)−C(8)−P(1) five-membered rings are essentially coplanar to within ±0.046 Å (for 3a) and ±0.104 Å (for 3b). The interplanar angles for 3a and 3b, as defined by the Se(1)/C(7)/C(8)/P(1) vs. Se(1)/M(1)/P(1) planes, are 4.9° and 11.4° respectively indicating marginal folding along the P(1)···Se(1) vector. In addition, there is a clear change in dihedral angle between the phenyl/phenylene rings attached to Se(1) as a function of chelation. In the free ligand 2a, the dihedral angle is 54.1° which expands to 87.0° (for 3a) and 82.1° (for 3b).

As a consequence, weak non covalent M(1)···Se(1) intermolecular contacts [Pt(1)···Se(1) 3.709 Å for 3a; Pd(1)···Se(1) 3.787 Å for 3b] link adjacent molecules into dimer pairs. Although not shown in Fig. 2, additional Se(1)···Cl(1A) interactions are present (3.445 Å for 3a; 3.679 Å for 3b). In the case of 3a this distance is significantly less than the sum of the van der Waals radii for Se and Cl atoms (3.65 Å).

Fig. 2 Molecular structures of (a) 3a and (b) 3b. All C−H hydrogen atoms have been omitted. Selected bond distances (Å) and angles (deg) for 3a: Pt(1)−P(1) 2.2094(11), Pt(1)−Se(1) 2.3561(5), Pt(1)−Cl(1) 2.3668(11), Pt(1)−Cl(2) 2.3142(11); P(1)−Pt(1)−Se(1) 90.28(3), P(1)−Pt(1)−Cl(2) 91.89(4), Se(1)−Pt(1)−Cl(1) 85.94(3), Cl(1)−Pt(1)−Cl(2) 92.04(4). For 3b: Pd(1)−P(1) 2.2615(4), Pd(1)−Se(1) 2.3656(2), Pd(1)−Cl(1) 2.3506(4), Pd(1)−Cl(2) 2.3180(5); P(1)−Pd(1)−Se(1) 87.902(12), P(1)−Pd(1)−Cl(1) 98.800(17), Se(1)−Pd(1)−Cl(1) 84.577(14), Cl(1)−Pd(1)−Cl(2) 88.709(18).

Having successfully demonstrated the preparation of aryl based selenoether phosphines, we also found the ease by which an alkyl backboned selenoether phosphine can be readily obtained. Using a AIBN radical initiated hydroselenation, reaction of Ph₂PCH=CH₂ with PhSeH afforded solid 4 (Scheme 2) regioselectively which has been characterised by the δ(P) = −15.4 ppm. Reaction of the anti-Markovnikov product 4 with PdCl₂(cod) gave the five-membered chelate complex 5 displaying the expected spectroscopic and analytical properties. The significant downfield shift in the ³¹P{¹H} NMR spectrum of 5 [δ(P) 62.3 ppm] is fully consistent with coordination of both P and Se donor atoms to a single Pd(II) metal centre.

In summary, we have shown how established synthetic routes can be used to access, hitherto unknown, selenoether modified tertiary phosphines and demonstrated their ease of P,Se-chelation at soft metal centres. Further coordination studies are in progress and results from these will be reported in due course.

Acknowledgements

We would like to thank the EPSRC and Loughborough University for funding (PMS). Johnson Matthey are gratefully acknowledged for the generous loan of precious metal salts.

Notes and References

† See ESI for synthetic details and characterising data for all new compounds.

Crystal data: For 2b, C₃₂H₂₅O₃PSe: Mr = 447.35, orthorhombic, space group P₂₁2₁2₁, a = 8.1752(5) Å, b = 10.1189(6) Å, c = 24.5004(14) Å, V = 2026.8(2) Å³, T = 150(2) K, Z = 4, μ(Mo-Kα) = 0.71073 Å, 17792 data measured, 4890 unique (Rint = 0.0366), dcalc = 1.466 g cm⁻³, R1 = 0.0391 (for 4049 data with I > 2σ(I)), wR2 = 0.0858 (all data), and 248 parameters. CCDC xxxxxx.

For 3a, C₃₂H₂₅P₃PPSe: Mr = 683.31, monoclinic, space group P2₁/n, a = 11.2539(7) Å, b = 13.4809(8) Å, c = 14.7665(9) Å, β = 101.0079(10)°, V = 2199.0(2) Å³, T = 150(2) K, Z = 4, μ(Mo-Kα) = 0.71073 Å, 19234 data measured, 5350 unique (Rint = 0.0357), dcalc = 2.064 g cm⁻³, R1 = 0.0276 (for 4130 data with I > 2σ(I)), wR2 = 0.0557 (all data), and 262 parameters. CCDC xxxxxx.

For 3b, C₃₂H₂₅Cl₃P₃PPSe: Mr = 709.58, triclinic, space group P T, a = 11.0639(5) Å, b = 11.0640(5) Å, c = 11.7585(5) Å, α = 89.7421(7)°, β = 76.3793(7)°, γ = 68.4046(6)°, V = 1298.32(10) Å³, T = 150(2) K, Z = 2, μ(Mo-Kα) = 0.71073 Å, 11577 data measured, 6012 unique (Rint = 0.0110), dcalc = 1.815 g cm⁻³, R1 = 0.0188 (for 5490 data with I > 2σ(I)), wR2 = 0.0464 (all data), and 302 parameters. CCDC xxxxxx. All three structures were determined routinely.


6 (a) V. Cadierno, J. Díez, J. García-Álvarez and J. Gimeno, Dalton Trans., 2010, 39, 941; (b) C. Kling, H. Ott, G. Schwab and D. Stalke, Organometallics, 2008, 27, 5038.


