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Citation: BOZIC-WEBER, B., 2013. Improving the photoresponse of copper(i) dyes in dye-sensitized solar cells by tuning ancillary and anchoring ligand modules. Dalton Transactions, 42(34), pp. 12293-12308.

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Metadata Record: https://dspace.lboro.ac.uk/2134/25424

Version: Published

Publisher: © The Royal Society of Chemistry

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Improving the photoresponse of copper(Ⅰ) dyes in dye-sensitized solar cells by tuning ancillary and anchoring ligand modules†

Biljana Bozic-Weber, Sven Y. Brauchli, Edwin C. Constable,*, Sebastian O. Fürer, Catherine E. Housecroft,*, Frederik J. Malzner, Iain A. Wright and Jennifer A. Zampese

The syntheses of five homoleptic copper(Ⅰ) complexes [CuL2][PF6] are described in which L is a 4,4’-di(4-bromophenyl)-6,6’-dialkyl-2,2’-bipyridine ligand (compounds 1–4 with methyl, ethyl, 2-ethylhexyl substituents, respectively) or 4,4’-di(4-bromophenyl)-6,6’-diphenyl-2,2’-bipyridine (5). The new ligands 2–5 and copper(Ⅰ) complexes [CuL2][PF6] (L = 1–5) have been fully characterized. The single crystal structures of 2{[Cu(1)2][PF6]}·3Me2CO, [Cu(2)2][PF6], 2{[Cu(3)2][PF6]}·Et2O and [Cu(5)2][PF6]·CH2Cl2 have been determined. The first three structures show similar distorted tetrahedral environments for the Cu+ centres with angles between the least squares planes of the bpy domains of 85.6, 86.4 and 82.9°, respectively; in contrast, the Cu+ ion in [Cu(5)2][PF6]·CH2Cl2 is in a flattened coordinate environment due to intra-cation face-to-face π-interactions. The solution absorption spectra of the complexes with ligands 1–4 are virtually identical with an MLCT band with values of \( \lambda_{\text{max}} \approx 481–488 \text{ nm} \). In contrast, the absorption spectrum of [Cu(5)2][PF6] shows two broad bands in the visible region. Cyclic voltammetric data show that oxidation of the copper(Ⅰ) centre occurs at a more positive potential in [Cu(2)2][PF6], [Cu(3)2][PF6] and [Cu(4)2][PF6] than in [Cu(1)2][PF6] or [Cu(5)2][PF6] with the latter being oxidized at the lowest potential. The complexes have been used to prepare dye-sensitized solar cells (DSCs) incorporating heteroleptic dyes of type [Cu(L)(Lmodified)]+ where L is a 4,4’-dimethyl-2,2’-bipyridine functionalized in the 4- and 4’-positions with phosphonic acid groups with (Lmodified = 7) and without (Lmodified = 6) a spacer between the metal-binding and anchoring domains. The presence of the spacer results in enhanced performances of the dyes, and the highest energy conversion efficiencies are observed for the dyes [Cu(3)(7)]+ (\( \eta = 2.43% \) compared to 5.96% for standard dye N719) and [Cu(5)(7)]+ (\( \eta = 2.89% \) compared to 5.96% for N719). Measurements taken periodically over the course of a week indicate that the cells undergo a ripening process (most clearly seen for [Cu(5)(6)]+ and [Cu(5)(7)]+) before their optimum performances are achieved. IPCE (EQE) data are presented and confirm that, although the photo-to-current conversions are promising (37–49% for \( \lambda_{\text{max}} \approx 480 \text{ nm} \)), the copper(Ⅰ) dyes do not realize the broad spectral response exhibited by N719.

Introduction

There is currently significant interest in the use of earth-abundant metals in dyes in dye-sensitized solar cells (DSCs). Sauvage and co-workers1,2 were the first to use copper(Ⅰ) complexes as sensitizers. The rationale for their application is based on similarities between the photophysical properties of copper(Ⅰ) and ruthenium(Ⅱ) complexes.3,4 The design of suitable ligands for copper(Ⅰ)-based dyes depends critically on information gained from photophysical and theoretical studies. The suitability of oligopyridine copper(Ⅰ)-based complexes as sensitizers has been confirmed by the results of TD-DFT calculations.4,5 In addition to determining orbital composition (in particular, confirming whether the LUMO of a complex possesses contributions from the anchoring units as required for efficient electron injection), DFT and TD-DFT calculations provide insight into the consequences of changes in ligand conformations upon excitation of the complex,6 probe energy matching of the HOMO and LUMO of the dye to the

Department of Chemistry, University of Basel, Spitalstrasse 51, CH4056 Basel, Switzerland. E-mail: edwin.constable@unibas.ch, catherine.housecroft@unibas.ch; Fax: +41 61 267 1018; Tel: +41 61 267 1008
†Electronic supplementary information (ESI) available: Fig. S1. Packing interaction in 2{[Cu(1)2][PF6]}·3Me2CO. CCDC 942069–942073. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c3dt51416a
conduction band of the semiconductor and predict absorption spectra of dyes.9–11

Our recent investigations of copper(II)-based dyes have capitalized on the lability of copper(II) bis(diimine) complexes in solution which allows heteroleptic complexes containing [Cu(bpy)2] cores (bpy = 2,2′-bipyridine) to be assembled in situ by facile ligand exchange.9,12–14 One of the two ligands in the complex is selected for its ability to anchor the complex to a mesoporous TiO2 surface. Typical choices of anchoring units are carboxylic acid, phosphonic acid or phenol groups, but we have recently shown that DSCs incorporating dyes with phosphonic acid anchoring groups typically exhibit the best performances,9,13,14 consistent with the observations of Grätzel.15

The ancillary ligand in the dye is designed to provide efficient light-harvesting properties. It is well established that a 6,6′-dimethyl substitution pattern is required in bpy ligands for stabilization of copper(I).16 On the other hand, the photophysical and electrochemical properties of copper(II) bis(diimine) complexes are strongly influenced by the electronic and steric effects of the substituents, viz. 6,6′-substituents in bpy and 2,9-substituents in 1,10-phenanthroline (phen).8,17–22

Here we investigate the effects of varying the steric demands of the 6,6′-substituents in the bpy derived ancillary ligands 1–4 (Scheme 1) and the effects of going from alkyl to phenyl substituents (5, Scheme 1), in conjunction with the effects of introducing an aromatic spacer in the phosphonic acid anchoring ligands 6 and 7 (Scheme 1). The inclusion of the bromo-substituent in ancillary ligands 1–5 provides a built-in handle for further derivatization and tuning of electronic properties of the ligand as we have illustrated with the decoration of 1 with first and second generation hole-transport dendrons.14 The choice of 6,6′-substituents in 1–4 was made to explore the effects of alkyl chain length and steric demands on the performance of the surface-anchored dye. In anticipation that longer chains might inhibit assembly of dye complexes on the surface, we decided to trial a new anchoring ligand, 7, in which a phenyl spacer extends the spatial separation of the metal centre from the anchoring units compared to dyes incorporating anchoring ligand 6 (Scheme 2).

Experimental

General

1H and 13C NMR spectra were recorded using a Bruker Avance III-250, 400 or 500 NMR spectrometer with chemical shifts referenced to residual solvent peaks with respect to δ(TMS) = 0 ppm. Absorption spectra were recorded on a Cary-5000 spectrophotometer, and FT-IR spectra on a Shimadzu 8400S instrument. ESI mass spectra were recorded on a Bruker Esquire 3000 plus instrument.

Electrochemical measurements were made on a CH Instruments 900B potentiostat using glassy carbon, platinum wire and silver wire as the working, counter, and reference electrodes, respectively. Samples were dissolved in HPLC grade CH2Cl2 (10−4 to 10−5 mol dm−3) containing 0.1 mol dm−3 [′Bu,N][PF6] as the supporting electrolyte; all solutions were degassed with argon. Cp2Fe was used as the internal reference.

(1E,5E)-1,6-Bis(4-bromophenyl)hexa-1,5-diene-3,4-dione and compound 1 were prepared as previously reported.14 Compound 6 was prepared by a modified literature method.24

1-(2-Oxohexyl)pyridinium iodide

Iodine (25.6 g, 101 mmol) was dissolved in pyridine under vigorous stirring (40.8 mL, 39.9 g, 505 mmol) and then hex-2-one (13 mL, 101 mmol) was added dropwise over a period of 5 min. The reaction mixture was heated at reflux overnight, and then the solvent was removed under vacuum. The crude material was used without purification for the preparation of compound 2. 1H NMR (400 MHz, CDCl3) δ/ppm 9.17 (d, J = 5.5 Hz, 2H, H9), 8.59 (m, 1H, H7), 8.06 (m, 2H, H29), 6.52 (s, 2H, HNCH3), 2.85 (t, J = 7.4 Hz, 2H, HCH2), 1.68 (m, 2H, HCH2), 1.38 (m, 2H, HCH2), 0.93 (t, J = 7.3 Hz, 2H, HMe).
Compound 2

1-(2-Oxohexyl)pyridinium iodide (3.00 g, 9.83 mmol) was dissolved in EtOH (60 mL) under vigorous stirring. (1E,5E)-1,6-Bis-(4-bromophenyl)hexa-1,5-diene-3,4-dione (1.03 g, 2.46 mmol) and NH₄OAc (2.84 g, 36.9 mmol) were added followed by EtOH (30 mL). The reaction mixture was heated at reflux for 2 days and was then allowed to cool to room temperature while being stirred. The precipitate was collected by filtration, washed with cold EtOH, and then dried under a stream of air. The product was recrystallized twice from EtOH, then from MeOH. Compound 2 was isolated as an off-white solid (0.630 g, 1.09 mmol, 44.3%). m.p. 157 °C. 1H NMR (500 MHz, CDCl₃) δ/ppm 8.48 (d, J = 1.6 Hz, 2H, H²), 7.63 (m, 8H, H²/B₂/R₃), 7.34 (d, J = 1.7 Hz, 2H, H⁵), 2.94 (t, J = 7.8 Hz, 4H, H¹), 1.83 (m, 4H, H⁴), 1.47 (m, 4H, H⁵), 0.99 (t, J = 7.4 Hz, 6H, H₃). 13C NMR (126 MHz, CDCl₃) δ/ppm 162.8 (CA₆), 156.7 (CA₂), 148.3 (CA₄), 138.1 (CB₁), 132.3 (CB₂/R₃), 123.3 (CA₄), 120.4 (CA₂), 116.6 (CA₃), 38.7 (CB₁), 32.2 (CB₁), 22.7 (C₇), 14.2 (CD). IR (KBr) ν/cm⁻¹ 3082 (w), 3053 (w), 2949 (m), 2924 (m), 2866 (w), 2853 (m), 2834 (w), 1583 (s), 1439 (m), 1385 (m), 1326 (m), 1165 (m), 1010 (s), 896 (m), 828 (s), 814 (s), 765 (s), 744 (m), 658 (m), 590 (m), 526 (s). ESI MS m/z 579.1 [M + H⁺] (calc. 579.1). UV-VIS (CH₂Cl₂, 1.0 × 10⁻³ mol dm⁻³) λ/nm 256 (ε/dm³ mol⁻¹ cm⁻¹ 58 350), 304 (14.360). Found: C 62.03, H 5.11, N 5.16; C₃₀H₃₀Br₂N₂ requires C 62.58, H 6.18, N 4.29%. This journal is © The Royal Society of Chemistry 2013

Compound 3

1-(4-Methyl-2-oxypropyl)pyridinium iodide

Iodine (30.4 g, 120 mmol) was dissolved in pyridine under vigorous stirring (248.5 mL, 47.4 g, 599 mmol) and then 4-methylpentan-2-one (15.0 mL, 120 mmol) was added dropwise over a period of 5 min. The reaction mixture was heated at reflux overnight. The reaction mixture was cooled to room temperature and then the solvent was removed under reduced pressure. The crude material was used without purification for the synthesis of compound 3. 1H NMR (400 MHz, CDCl₃) δ/ppm 9.22 (d, J = 5.7 Hz, 2H, H²/P₃), 8.58 (t, J = 7.8 Hz, 1H, H²/P₅), 8.09 (m, 2H, H²/P₅), 6.38 (s, 2H, H²/P₅), 2.70 (d, J = 6.8 Hz, 2H, H²/P₃), 2.20 (m, 1H, H²/P₃), 0.93 (d, J = 6.7 Hz, 6H, H²/M₄). 13C NMR (126 MHz, CDCl₃) δ/ppm 162.8 (CA₆), 156.7 (CA₂), 148.3 (CA₄), 138.1 (CB₁), 132.3 (CB₂/R₃), 123.3 (CA₄), 120.4 (CA₂), 116.6 (CA₃), 38.6 (CB₁), 32.9 (CB₁), 22.7 (C₇), 14.2 (CD). IR (KBr) ν/cm⁻¹ 3082 (w), 3053 (w), 2949 (m), 2924 (m), 2866 (w), 2853 (m), 2834 (w), 1583 (s), 1439 (m), 1385 (m), 1326 (m), 1165 (m), 1010 (s), 896 (m), 828 (s), 814 (s), 765 (s), 744 (m), 658 (m), 590 (m), 526 (s). ESI MS m/z 579.1 [M + H⁺] (calc. 579.1). UV-VIS (CH₂Cl₂, 1.0 × 10⁻³ mol dm⁻³) λ/nm 256 (ε/dm³ mol⁻¹ cm⁻¹ 58 350), 304 (14.360). Found: C 62.03, H 5.11, N 5.16; C₃₀H₃₀Br₂N₂ requires C 62.58, H 6.18, N 4.29%.

Compound 4

1-(2-Oxoctyl)pyridinium iodide

Iodide (16.2 g, 63.9 mmol) was dissolved in pyridine under vigorous stirring (25.8 mL, 25.2 g, 319 mmol), and then octan-2-one (10.0 mL, 63.9 mmol) was added dropwise over a period of 5 min. The reaction mixture was heated to reflux overnight. The reaction mixture was cooled to room temperature and then the solvent was removed under reduced pressure. The crude material was used without purification for the synthesis of compound 4. 1H NMR (500 MHz, CDCl₃) δ/ppm 9.24 (d, J = 5.8 Hz, 2H, H²/P₃), 8.54 (m, 1H, H²/P₅), 8.07 (m, 2H, H²/P₅), 6.47 (s, 2H, H²/P₅), 2.82 (t, J = 7.4 Hz, 2H, H²/P₅), 1.65 (m, 2H, H²/P₅), 1.27 (overlapping m, 6H, H²/P₅/CH₃/CH₂/CH₂), 0.85 (t, J = 6.5 Hz, 2H, H²/M₄).
1.3 Hz, 2H, H_{Py^2}), 8.65 (m, 1H, H_{Py^4}), 8.17–8.08 (m, 4H, H_{Py^2+3,Py^3}), 7.78 (m, 1H, H_{Ph^4}), 7.64 (m, 2H, H_{Ph^2}), 6.4 (s, 2H, H_{Ph^CH^2}).

**Compound 5**

Compound 5 was prepared in the same manner as 2 starting with 1-(2-phenyl-2-oxoethyl)pyridinium iodide (1.61 g, 4.95 mmol) and (1E,5E)-1,6-bis(4-bromophenyl)hexa-1,5-diene-3,4-dione (1.04 g, 2.48 mmol) and NH$_4$OAc (2.68 g, 4.95 mmol). Single crystals were grown by diffusion of pentane and hexane (802 mg, 65%).

**Compound 7a**

Compound 7a (200 mg, 0.33 mmol) was dissolved in aqueous HCl (6 M, 20 mL) and the solution was heated at reflux for 48 h. The solvent was removed under reduced pressure to leave a pale yellow residue. The residue was treated with AcOH (15 mL) at reflux in the presence of 5 drops of concentrated aqueous HCl prior to addition of water. This resulted in the precipitation of 7 as a white powder (163 mg, >99%), m.p. >300 °C. 1H NMR (500 MHz, DMSO-d$_6$) δ/ppm 8.35 (s, 2H, H_{H^4}), 7.95 (dd, J$_{HH} = 1.5$ Hz, J$_{PH} = 7.7$ Hz, 8H, HB2/B3), 2.37 (s, 6H, HMe). 13C NMR (126 MHz, DMSO-d$_6$) δ/ppm 159.7 (C_{A2}), 135.9 (C_{A4/B4}), 134.4 (C_{A4/B4}), 131.3 (C_{B1}), 126.5 (C_{B5}), 121.2 (C_{A5}), 115.2 (C_{A5}), 24.1 (C_{A5}), C_{A5} not resolved. IR (solid, ν/cm$^{-1}$) 3424 (w), 2968 (w), 1606 (m), 1597 (s), 1539 (m), 1538 (m), 1510 (m), 1491 (m), 1457 (w), 1403 (m), 1385 (w), 1336 (w), 1320 (m), 1275 (m), 1238 (s), 1178 (s), 1154 (m), 1017 (s), 978 (m), 937 (s), 921 (s), 880 (s), 826 (vs), 761 (m), 728 (w), 680 (vs), 564 (s). Found C 55.06, H 4.96, N 5.46; C$_{24}$H$_{22}$N$_2$O$_6$P$_2$·1.5H$_2$O requires C 55.07, H 4.81, N 5.35%. Protonated 7, see text: 1H NMR (500 MHz, CD$_2$CO$_2$D) δ/ppm 8.66 (s, 2H, H_{A^4}), 8.33 (s, 2H, H_{B^2}), 8.15 (dd, J$_{HH} = 7.6$ Hz, J$_{PH} = 14.0$ Hz, 8H, HB2/B3), 3.29 (s, 2H, H_{Me}). 13C NMR (126 MHz, CD$_2$CO$_2$D) δ/ppm 160.8 (C_{A^4}), 144.2 (C_{A^2}), 140.2 (d, J$_{PC} = 3.2$ Hz, C_{B^1}), 135.2 (d, J$_{PC} = 11.4$ Hz, C_{B^1}), 134.3 (d, J$_{PC} = 187$ Hz, C_{B^1}), 130.6 (d, J$_{PC} = 16.2$ Hz, C_{B^2}), 130.2 (C_{C^3}), 126.7 (C_{A^3}), 21.2 (C_{Me}).

**[Cu$_2$]([PF$_6$]) general procedure**

In a typical reaction, [Cu(NCMe)$_4$][PF$_6$] was dissolved in a mixture of MeCN and CH$_2$Cl$_2$ and ligand was added. The solution turned deep red immediately. It was stirred overnight, after which time the volume of solvent was reduced in vacuo and the product was precipitated by addition of Et$_2$O. The solid was redissolved in MeCN and precipitated again by addition of Et$_2$O. The dark red product was collected by filtration and was washed with cold Et$_2$O and dried in a stream of air.

**[Cu(1)]([PF$_6$])**

[Cu(1)]([PF$_6$]) (85.0 mg, 0.228 mmol) was reacted with 1 (225 mg, 0.456 mmol) in MeCN (4 mL) and CH$_2$Cl$_2$ (16 mL). [Cu(1)]([PF$_6$]) was isolated as a dark red solid (217 mg, 0.181 mmol, 79.6%). Decomp. > 320 °C. 1H NMR (500 MHz, CD$_2$CN) δ/ppm 8.66 (br, 4H, H_{H^4}), 7.88 (dd, J$_{PH} = 7.9$ Hz, 8H, H$_{B2/B3}$), 7.83 (br, 4H, H$_{B2}$), 7.78 (d, J = 7.7 Hz, 8H, H$_{B2/B3}$), 2.37 (s, 2H, H$_{2}$), 13C NMR (126 MHz, CD$_2$CN) δ/ppm 161.5 (C_{A^4}), 153.7 (C_{A^2}), 149.8 (C_{A^3}), 137.1 (C_{B^1}), 133.3 (C_{B2/B3}), 130.2
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(C(82Br3)), 124.5 (C(84)), 124.8 (C(85)), 118.7 (C(86)), 25.1 (C(8)). ESI MS m/z 1050.8 [M – PF4]− (calc. 1050.9), 495.0 [1+H]+ (calc. 495.0). UV–VIS (CH2Cl2, 1.0 × 10−5 mol dm−3) λ/nm 280 (ε/dm3 mol−1 cm−1 82 600), 324 (sh) (42 200), 488 (11 300). Found: C 47.55, H 3.16, N 4.93; C68H36Br4CuF6N4P·H2O requires C 55.32, H 5.21, N 4.18. C66H72Br4CuFe2N2P2MeCN requires C 55.32, H 5.21, N 4.21%.

[Cu(2)][PF6]

[Cu(NCMe)4][PF6] (58.7 mg, 0.158 mmol) was reacted with 0.147 mmol) in MeCN (10 mL) and CH2Cl2 (10 mL). [Cu(2)][PF6] was isolated as a dark red solid (296 mg, 0.143 mmol, 58.9%). m.p. 325 °C. 1H NMR (500 MHz, CDCl3) δ/ppm 8.67 (broadened d, 4H, HA3), 7.89 (d, J = 6.8 Hz, Hb), 2.68 (s, 8H, Ha), 1.41 (s, 8H, Hb), 0.89 (s, 8H, Hc), 0.72 (s, 8H, Hd), 0.55 (t, J = 6.6 Hz, 12H, H4). 13C NMR (126 MHz, CDCl3) δ/ppm 163.2 (C(82)), 153.6 (C(82)), 150.1 (C(82)), 137.0 (C(83)), 133.3 (C(84)), 130.2 (C(85)), 124.9 (C(86)), 124.1 (C(87)), 118.9 (C(88)), 40.7 (C(8)), 32.0 (C(9)), 30.0 (C(10)), 22.9 (C(11)), 14.0 (C(12)). ESI MS m/z 1331.0 [M – PF4]− (calc. 1331.2), 635.2 [5+H]2+ (calc. 635.1). UV–VIS (CH2Cl2, 1.0 × 10−5 mol dm−3) λ/nm 282 (ε/dm3 mol−1 cm−1 85 000), 325 sh (41 100), 486 (10 800). Found: C 55.44, H 5.21, N 4.18; C66H72Br4CuFe2N2P2MeCN requires C 55.32, H 5.21, N 4.18.

[Cu(2)][PF6]

[Cu(NCMe)4][PF6] (90.4 mg, 0.243 mmol) was reacted with 5 (300 mg, 0.485 mmol) in MeCN (5 mL) and CH2Cl2 (40 mL). [Cu(5)][PF4] was isolated as a dark red solid (206 mg, 0.143 mmol, 58.9%). m.p. 325 °C. 1H NMR (500 MHz, CDCl3) δ/ppm 8.23 (d, J = 1.2 Hz, 2H, H4), 7.77 (overlapping m, 8H, H8-B8), 7.64 (d, J = 1.2 Hz, 2H, H6), 7.62 (d, J = 7.5 Hz, 4H, H8), 7.05 (t, J = 7.4 Hz, 2H, H7). 13C NMR (126 MHz, CDCl3) δ/ppm 156.9 (C(86)), 153.6 (C(84)), 149.5 (C(83)), 137.9 (C(81)), 135.6 (C(84)), 132.9 (C(82)), 128.9 (C(83)), 127.7 (C(81)), 127.6 (C(82)), 124.8 (C(84)), 119.4 (C(85)). ESI MS m/z 1299.2 [M – PF4]− (calc. 1399.0), 618.8 [5 + H]+ (calc. 619.0). UV–VIS (CH2Cl2, 1.0 × 10−5 mol dm−3) ε/nm 245 sh (48 000), 268 (72 400), 289 (76 000), 335 (29 300), 423 (5840), 573 (4530). Found: C 55.02, H 3.20, N 4.14; C60H60Br4CuF6N4P requires C 55.14, H 3.27, N 3.78%.

Crystallography

Data were collected on a Bruker APEX-II diffractometer with data reduction, solution and refinement using the programs APEX2 and SHELXL97 or SHELX-13.28 ORTEP-type diagrams and structure analysis used Mercury v. 3.0,27,28
2\([\text{Cu}(3)_2[\text{PF}_6]]\)_2\cdot\text{Et}_2\text{O}

\[\text{C}_{124}\text{H}_{130}\text{Br}_8\text{Cu}_2\text{F}_{12}\text{N}_8\text{O}_2\text{P}_2, \text{M} = 2804.61, \text{yellow plate, monoclinic, space group } \text{P2}_1/c, a = 15.5265(7), b = 23.9138(9), c = 17.7749(7) \text{ Å}, \beta = 114.893(2)^\circ, U = 5976.64(4) \text{ Å}^3, Z = 2, D_c = 1.556 \text{ Mg m}^{-3}, \mu(\text{Cu-Ka}) = 4.410 \text{ mm}^{-1}, T = 123 \text{ K}. \text{Total 46 225 reflections, 10 612 unique, } R_{\text{int}} = 0.0480. \text{Refinement of 8760 reflections (761 parameters) with } I > 2\sigma(I) \text{ converged at final } R_1 = 0.0467 (R_F \text{ all data} = 0.0581), wR_2 = 0.1168 (wR_F \text{ all data} = 0.1244), gof = 1.037. \text{CCDC 942071.}

\[\text{[Cu}(5)_3[\text{PF}_6]\text{]}\cdot\text{CH}_2\text{Cl}_2\]

\[\text{C}_{69}\text{H}_{46}\text{Br}_4\text{Cl}_2\text{CuF}_6\text{N}_4\text{P}, \text{M} = 1530.12, \text{yellow plate, orthorhombic, space group } \text{Pbcn, } a = 7.53560(15), b = 28.3927(6) \text{ Å}, U = 5976.2(2) \text{ Å}^3, Z = 4, D_c = 1.701 \text{ Mg m}^{-3}, \mu(\text{Cu-Ka}) = 5.284 \text{ mm}^{-1}, T = 123 \text{ K}. \text{Total 37 392 reflections, 5233 unique, } R_{\text{int}} = 0.318. \text{Refinement of 1637 reflections (395 parameters) with } I > 2\sigma(I) \text{ converged at final } R_1 = 0.0825 (R_F \text{ all data} = 0.2739), wR_2 = 0.1280 (wR_F \text{ all data} = 0.1881), gof = 0.933. \text{CCDC 942073.}

Solar cell fabrication

The fabrication of DSCs was based on the method of Grätzel and coworkers.\(^{29}\) The TiO\(_2\) paste was prepared adapting the published method using a three-roll mill (50 EC, EXAKT, Germany), a sonicator bath and terpineol (CAS: 8000-41-7). Each working electrode was made from an FTO glass plate (Solaronix TCO22-7, 2.2 mm thickness, sheet resistance \(\approx 7 \Omega\) square\(^{-1}\) which was cleaned by sonicating in Hellmanex\(^\text{®}\) surfactant (2% in milliQ water), and rinsed with milliQ water and EtOH. After surface activation in a UV-O\(_3\) system (Model 235-220, Jelight Company Inc.) for 20 min, the FTO plate was immersed in aqueous TiCl\(_4\) (40 mmol dm\(^{-3}\)) at 70 °C for 30 min, and then washed with milliQ water and EtOH. The FTO plate was dried and a layer of TiO\(_2\) paste was screen printed (90 T, Serilith AG, Switzerland). The printed plate was kept in an EtOH chamber for 3 min to reduce surface irregularities of the printed layer and dried for 6 min at 125 °C on a heating plate. The screen printing process was repeated 8 times and then the electrodes were gradually heated at 75 °C for 30 min, at 135 °C for 15 min, at 325 °C for 5 min, at 375 °C for 5 min, at 450 °C for 15 min and at 500 °C for 15 min. After the final sintering, the thickness of the TiO\(_2\) layer was 14–16 μm (measured with a Tencor Alpha-Step 500 profilometer). The annealed TiO\(_2\) film was post-treated with 40 mmol dm\(^{-3}\) aqueous TiCl\(_4\) solution (see above), rinsed with milliQ water and EtOH and sintered at 500 °C for 30 min.

The electrodes were cooled to ca. 80 °C and immersed in a 1 mM DMSO solution of the anchoring ligand 6 or 7 for 20 h. The colourless electrode was removed from the solution, washed with DMSO and EtOH and dried under a stream of N\(_2\). The electrode with the adsorbed anchoring ligand was immersed in a 1 mM MeCN solution of [Cu(L\(_2\))\(_2\)[PF\(_6\)] (L = 1–4) and in a 1 mM CH\(_2\)Cl\(_2\) solution of [Cu(5)\(_3\)][PF\(_6\)] for 4 days to produce red-orange coloured electrodes. The electrodes were removed from the solution and were washed with CH\(_2\)Cl\(_2\) and dried under a stream of N\(_2\).

Each counter electrode was prepared from an FTO glass plate (Solaronix TCO22-7, 2.2 mm thickness, sheet resistance \(\approx 7 \Omega\) square\(^{-1}\)) with a previously drilled hole. Residual organic impurities were removed by heating for 15 min at 500 °C on a heating plate and the perforated plate was washed with water, then 0.1 M HCl solution in EtOH and finally ultrasonicated in an acetone bath for 15 min. It was dried on a heating plate at 500 °C for 15 min. The Pt catalyst was deposited on the FTO glass plate by coating with Platsisol T (2 \(\times\) 25.0 μL), Solaronix, and dried on a heating plate at 500 °C for 15 min.

The dye-covered TiO\(_2\) electrode and Pt counter-electrode were assembled using a thermoplas hot-melt sealing foil (Solaronix, Meltonix 1170-25 Series, 25 microns thick) by heating while pressing them together. The electrolyte comprised LiI (0.1 mol dm\(^{-3}\)), I\(_2\) (0.05 mol dm\(^{-3}\)), 1-methylbenzimidazole (0.5 mol dm\(^{-3}\)) and 1-butyl-3-methylimidazolium iodide (0.6 mol dm\(^{-3}\)) in methoxypropionitrile, and was introduced into the cell by vacuum backfilling. The hole on the counter electrode was finally sealed using the hot-melt sealing foil and a cover glass.

The solar cell measurements and testing protocol were performed using fully masked cells. A black coloured copper sheet was used for masking with a single aperture of an average area of 0.06012 cm\(^2\) (with a standard deviation of 1%) placed over the screen printed dye-sensitized TiO\(_2\) circle. The area of the aperture in the mask was smaller than the active area of the dye-sensitized TiO\(_2\) dot (0.288 cm\(^2\)). For complete masking, tape was also applied over the edges and rear of the cell. Current density–voltage (\(I–V\)) measurements were made by irradiating from behind using a light source SolarSim 150 (100 mW cm\(^{-2}\) = 1 sun). The power of the simulated light was calibrated by using a reference Si photodiode. The standard dye N719 was purchased from Solaronix.

The quantum efficiency measurements were performed on an Spe-Quest quantum efficiency setup from Rera Systems (Netherlands) equipped with a 100 W halogen lamp (QTH) and a lambda 300 grating monochromator from Lot Oriel. The monochromatic light was modulated to 3 Hz using a chopper wheel from ThorLabs. The cell response was amplified with a large dynamic range IV converter from CVI Melles Griot and then measured with an SR830 DSP Lock-In amplifier from Stanford Research.

Results and discussion

Synthesis and characterization of ancillary ligands

We have previously described the synthesis of compound \(1\) using Krönkhe methodology\(^{30}\) by the treatment of \((\text{1E,3E})\)-1,6-bis(4-bromophenyl)hexa-1,5-diene-3,4-dione with 1-(2-oxo-propyl)pyridinium chloride and ammonium acetate.\(^{14}\) Compounds 2, 3, 4, and 5 were prepared according to Scheme 3 using the appropriate “pyridacylpypyridinium salts” to introduce
the desired R groups at the 6,6′-positions of the 2,2′-bipyridine.

The electrospray mass spectrum of each of compounds 2–5 exhibited a peak envelope as the base peak which corresponded to [M + H]+ and with an isotope pattern characteristic of the presence of two bromine atoms. The 1H and 13C NMR spectra were in accord with a 4,4′,6,6′-substituted 2,2′-bipyridine, and the spectra were assigned using COSY, NOESY, DEPT, HMOC and HMBC methods. The 1H NMR resonances for protons H4 and H1 were distinguished by the appearance of a cross peak in the HMBC spectrum of each of compounds 2–5 between signals for Cpp and H1. In the 500 MHz 1H NMR spectra of 2, 3 and 4, the signals for H8 and H9 overlap, and unambiguous assignments of the C8 and C9 resonances were not made from the HMOC or HMBC spectra.

Compounds 1–4 each contain alkyl substituents in the 6,6′-positions and their electronic absorption spectra are similar, being dominated by an intense band at 255 nm with a weaker band close to 305 nm (Fig. 1). Introduction of the 6- and 6′-phenyl substituents in compound 5 red-shifts both absorptions (to 265 and 321 nm). The enhancement of intensity of the higher energy band is consistent with the latter arising from π*→π transitions. The lower energy absorption shows no change in intensity along the series of five compounds, and is assigned to π*→n(py) transitions.

In terms of ligand design for dyes for DSCs, compound 5 appears superior to 1–4 in terms of red-shifting of the spectral response, and we later confirm that this trend is preserved in the homoleptic copper(i) complexes of these ligands.

**Synthesis and characterization of anchoring ligand 7**

Compatibility between anchoring and ancillary ligands is of prime importance in our copper-containing heteroleptic dyes. We have previously made extensive use of anchoring ligand 6 and have shown that, even though 6 binds the copper(i) ion close to the TiO2 surface, it remains suited to ancillary ligands with sterically demanding substituents. On the other hand, we argued that the in situ assembly of copper(i) dyes in which the bpy-derived ancillary ligand contains long chain or aromatic 6,6′-substituents could be enhanced by introducing a spacer between the metal-binding domain of the anchoring ligand and the phosphonic acid anchors. With this in mind, we prepared ligand 7 (Scheme 4) as the first in a series of new phosphonic acid anchoring ligands.

Scheme 4 summarizes the synthesis of 7 from the corresponding ester, 7a (in turn obtained from ligand 1); preparation of the latter was based upon a literature method.

![Scheme 3: Synthetic route to compounds 2–5.](image)

![Scheme 4: Synthesis of 7.](image)
deprotection by treatment with aqueous HCl and workup, 7a was isolated in 65% yield. The base peak in the ESI mass spectrum (m/z 609.3) arose from the [M + H]+ ion and a much less intense peak at m/z 631.2 was assigned to [M + Na]+.

X-ray quality crystals of 7a were grown by diffusion of a mixture of pentane and hexane (1:1 by volume) into a concentrated acetone solution of the compound. Structural determination confirmed the structure shown in Fig. 2. The bpy unit adopts the expected trans conformation and the two halves of the molecule are related by an inversion centre. Compound 7a is disordered and the OEt group containing O3 has been modelled over two sites with 49 and 51% occupancies; the methyl group of the second OEt unit has been modelled over two sites with 49 and 51% occupancies; the methyl substituents prevent edge-to-face stacking interactions. The dominant packing forces involve CH···O and CH···N contacts.

Quantitative conversion of the ester 7a into phosphonic acid 7 was achieved using aqueous HCl (6 M) at reflux followed by treatment of the residue obtained with aqueous acetic acid. This latter step appears to be essential for the isolation of pure 7. Compound 7 is poorly soluble in most common solvents and it was not possible to obtain an electrospray mass spectrum. However, satisfactory elemental analysis was obtained for the hydrate 7·H2O. The solubility of 7 in DMSO-d6 was sufficient to permit a 1H NMR spectrum to be recorded, but 13C NMR resonances could only be assigned using HMOC and HMBC methods; no cross peaks associated with C32 (see Scheme 1 for atom labelling) in the HMBC spectrum were observed. Compound 7 dissolved readily in CF3CO2D. The 1H NMR spectrum was consistent with protonation of 7, although it is difficult to assess from the change in the chemical shift of the signal for H31 (δ 8.53 ppm for 7 in DMSO-d6 to δ 8.67 ppm for the compound in CF3CO2D) whether the species present is [H27]2+ (trans-conformation of the bpy unit) or [H7]2+ (cis-bpy). These conformational changes have been confirmed structurally for salts of [Hdmmbpy]2+ and [H2dmbpy]2+ (dmbpy = 6,6′-dimethyl-2,2′-bipyridine),32 and we have recently observed that addition of aliquots of trifluoroacetic acid to a CDCl3 solution of a ligand closely related to 7 leads to an initial downfield shift of the signals for H31 from δ 8.39 to 8.53 ppm, followed by an upfield shift to δ 8.34 ppm corresponding to formation of first the monoprotonated and then the diprotonated ligand.33 Fully assigned NMR spectroscopic data for protonated 7 are given in the Experimental section.

Synthesis and characterization of homoletic copper(i) complexes

Reaction of each of ligands 1–5 with [Cu(NCMe)4]PF6 in CH2Cl2 at room temperature resulted in the formation of a red complex characterized as [CuL2]PF6. In the electrospray mass spectrum of each complex, the base peak corresponded to [M−PF6]−. The combination of copper and bromine isotopes leads to the two most intense peaks in the isotope pattern for the [M−PF6]− peak being of almost equal intensity. This is exemplified for [Cu(3)]2+PF6 in Fig. 3. The lower m/z value of each pair is listed in the Experimental section.

The 1H and 13C NMR spectra of the complexes were consistent with a single ligand environment in each complex. The spectra have been fully assigned using routine 2D methods.

X-ray quality crystals of 2[[Cu(1)]2PF6]-3Me2CO were grown by Et3O diffusion into an acetone solution of the complex at room temperature, and Fig. 4 shows the structure of the [Cu(1)]PF6 cation. The distorted tetrahedral coordination geometry of atom Cu1 is less flattened than in [Cu(dmbpy)2]+ (dmbpy = 6,6′-dimethyl-2,2′-bipyridine); in [Cu(1)]PF6, the angle between the least squares planes containing Cu1 and each bpy unit is 85.6°, compared to angles of 74.3° in [Cu(dmbpy)2]+, 80.9° in [Cu(dmbpy)2]BF4 and 80.9° in [Cu(dmbpy)2]ClO4.35 The flexibility of the Cu(bpy)2-unit is consistent with the suggestion35 that packing forces are responsible for the detailed geometry at the copper(i) centre. The two bpy units are close to planar (angles between the planes of the rings containing N1/N2 and N3/N4 = 5.5 and 7.9°, respectively), and the angles between the planes of the pairs of bonded phenyl and pyridine rings range from 10.1 to 34.9°. The variation in the latter is a consequence of packing interactions. Pairs of phenylpyridine units containing N1/C13 and N1′/C13′ (symmetry

![Fig. 3](image3.png)
code $i = 1 - x, -y, 1 - z$ sandwich one acetone solvent molecule which is disordered over two sites positioned about an inversion centre. The phenyl ring of the phenylpyridine unit containing N2 is disordered over two sites positioned about an inversion centre. The phenyl ring of the phenylpyridine unit containing N3 and N4 are disordered; each of the three terminal C atoms, C31, C32 and C33 with fractional occupancies 0.54 and 0.46, and C46, C47 and C48 with fractional occupancies 0.51 and 0.49. The presence of the disorder makes it meaningless to comment on the relative orientations of the butyl groups. Nonetheless, using the major occupancy sites, Fig. 5b illustrates how the butyl chains wrap around the copper(i) centre.

Single crystals of $[\text{Cu}(\text{2})]_{2}^{+}$ were grown by Et$_2$O diffusion into an acetonitrile solution of the complex; the structure of the $[\text{Cu}(\text{2})]_{2}^{+}$ cation is shown in Fig. 6. The distorted tetrahedral coordination geometry of Cu1 is essentially the same as that in $[\text{Cu}(\text{1})]_{2}^{+}$ with an angle between least squares planes containing Cu1 and each bpy unit of 86.4°. Each bpy unit is slightly twisted (angles between the planes of the rings containing N1/N2 and N3/N4 of 12.3°, and this is probably a consequence of steric crowding originating from the four isobutyl groups which are more twisted than in the complexes containing 1 and 2: angles between the planes of the rings with N1/N2 and N3/N4 of 12.3 and 29.0°, and this is probably a consequence of steric crowding originating from the four isobutyl groups which...
wrap around the metal centre (Fig. 6c). The isobutyl group containing C1 and C2 is disordered and has been modelled over two positions with site occupancies of 0.59 and 0.41; the carbon atom of the CH₂ group is common to both positions. The distortion of the ligand containing N3 and N4 extends to the bromophenyl unit with Br4; the phenylpyridine unit is significantly bowed. The origin of the distortion is not clear but may be associated with interactions with the ordered [PF₆]⁻ ion which is sandwiched between the phenyl rings containing C52 and C58 with three short CH···F contacts in the range 2.59 and 2.71 Å. In addition, atom Br4 exhibits a short contact with Br2′ (Br4···Br2′ = 3.6922(7) Å, symmetry code i = 1 − x, ½ + y, ½ − z, van der Waals radius of Br = 1.95 Å).36 As in the two previous structures, the extent that each phenylpyridine unit deviates from planarity is controlled by packing forces with twist angles lying in the range 20.1–40.3°. The Et₂O molecule is disordered and has been modelled over two sites on a special position.

Crystals of [Cu(5)]₂[PF₆]CH₂Cl₂ suitable for X-ray diffraction were grown by diffusion of Et₂O into a CH₂Cl₂ solution of the complex. Fig. 8 shows the structure of the [Cu(5)]₂⁺ cation and bond lengths and angles involving atom Cu1 are given in the caption. The complex crystallizes in the orthorhombic space group Pbcn with atom Cu1 lying on a two-fold axis. In contrast to [Cu(1)]₂⁺, [Cu(2)]₂⁺ and [Cu(3)]₂⁺, the [Cu(5)]₂⁺ cation possesses a flattened structure (Fig. 5) with an angle between the least squares planes containing Cu1 and each bpy unit of only 44.1°. This is a consequence of a π-stacking interaction between each 6- (or 6')-phenyl-substituent and a pyridine ring of the adjacent ligand. To achieve four such interactions, the bpy domains deviate from planarity (19.6 and 30.3°, respectively) and each phenyl ring twists (31.2 and 62.4°) with respect to the pyridine ring to which it is bonded. Despite these conformational changes, optimal π-stacking interactions with parallel rings are not achieved. The structural features of the copper(i) coordination sphere in [Cu(5)]⁺ are consistent with the flattening observed in a range of {[Cu(6,6'-Ph₂bpy)]}⁺, {[Cu(2,9-9-Ph₂phen)]}⁺ and related complexes,²⁰,²¹,³⁷–⁵⁰ although within these examples, the efficiency of the intra-cation π-interactions is variable. In [Cu(5)]₂[PF₆]CH₂Cl₂, face-to-face π-stacking occurs between pendant phenyl rings of neighbouring cations leading to columns of [Cu(5)]₂⁺ cations running parallel to the a-axis. Additional packing interactions involve short Br···Br contacts, and close CH···F and CH···Cl contacts; both the [PF₆]⁻ ion and the CH₂Cl₂ molecule are ordered.

The solution electronic absorption spectra of CH₂Cl₂ solutions (1.0 × 10⁻⁵ mol dm⁻³) of [CuL₂][PF₆] for L = 1–5 are shown in Fig. 9. Spectra for the complexes containing the aliphatic side chains are virtually superimposable with high energy bands assigned to ligand-based π→π transitions and a broad MLCT band in the visible region (λmax = 481–488 nm). The dual bands observed in the visible region of the spectrum [Cu(5)]₂[PF₆] and the red-shift to 573 nm are consistent with features in the absorption spectrum of [Cu(dpp)]⁺ (dpp = 2,9-diphenyl-1,10-phenanthroline) which
are explained in terms of the flattened, low-symmetry structure of the complex.9,20 The spectrum of [Cu(3)d][PF6] retains the enhanced spectral response in the higher energy bands observed for ligand 5 compared to 1–4 (Fig. 1).

All the complexes are electrochemically active and cyclic voltammetric data are presented in Table 1. It is well established that in copper(I) diimine complexes, the presence of substituents adjacent to the N,N-donor sites (e.g. 6,6'-positions in 2,2'-bpy or 2,9-sites in 1,10-phenanthroline) stabilizes copper(I) with respect to oxidation by sterically hindering the flattening of the coordination sphere that accompanies the d10 (Cu+) to d9 (Cu2+) electronic configurational change.19,22 Cyclic voltammograms were recorded in CH2Cl2 to avoid problems associated with coordinating solvents such as MeCN.19 The potentials for the metal-centred oxidation process in [CuL2][PF6] with L = 1–4 follow the expected trend, with the complex containing the least sterically demanding (methyl) substituents being the easiest to oxidize: [Cu(1)2]+ < [Cu(2)2]+ ≈ [Cu(3)2]+ ≈ [Cu(4)2]+. The oxidation potential for [Cu(1)2][PF6] (+0.42 V) compares to +0.50 V (vs. Fc/Fc′) in CH2Cl2 for [Cu(dmp)2][PF6] (dmp = 2,9-dimethyl-1,10-phenanthroline).

### Table 1

<table>
<thead>
<tr>
<th>Complex</th>
<th>Eox1/2/V</th>
<th>Evs/N</th>
<th>Ered1/2/V</th>
</tr>
</thead>
<tbody>
<tr>
<td>[Cu(1)2][PF6]</td>
<td>+0.42 (94)</td>
<td>+1.49</td>
<td>−2.16 (7)</td>
</tr>
<tr>
<td>[Cu(2)2][PF6]</td>
<td>+0.52 (87)</td>
<td>+1.47</td>
<td>−2.16 (9)</td>
</tr>
<tr>
<td>[Cu(3)2][PF6]</td>
<td>+0.54 (96)</td>
<td>+1.49</td>
<td>−2.20 (7)</td>
</tr>
<tr>
<td>[Cu(4)2][PF6]</td>
<td>+0.54 (84)</td>
<td>+1.49</td>
<td>−2.14 (7)</td>
</tr>
<tr>
<td>[Cu(5)2][PF6]</td>
<td>+0.36 (86)</td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

*No reduction process observed within the solvent accessible window.*

### Table 2

Sealed and masked DSC cell performances using anchoring ligands 6 and 7 with respect to standard dye N719 measured under the same conditions.

<table>
<thead>
<tr>
<th>Complex</th>
<th>On the day of sealing the cell</th>
<th>3 days after sealing the cell</th>
<th>4 days after sealing the cell</th>
</tr>
</thead>
<tbody>
<tr>
<td>[CuL2]+</td>
<td>Jsc/μA cm−2</td>
<td>Voc/V</td>
<td>ff</td>
</tr>
<tr>
<td>[Cu(1)2]</td>
<td>6</td>
<td>3.21</td>
<td>477</td>
</tr>
<tr>
<td>[Cu(2)2]</td>
<td>6</td>
<td>3.64</td>
<td>478</td>
</tr>
<tr>
<td>[Cu(3)2]</td>
<td>6</td>
<td>4.32</td>
<td>515</td>
</tr>
<tr>
<td>[Cu(4)2]</td>
<td>6</td>
<td>3.57</td>
<td>475</td>
</tr>
<tr>
<td>[Cu(5)2]</td>
<td>6</td>
<td>0.35</td>
<td>473</td>
</tr>
<tr>
<td>[Cu(1)2]</td>
<td>7</td>
<td>5.09</td>
<td>497</td>
</tr>
<tr>
<td>[Cu(2)2]</td>
<td>7</td>
<td>5.50</td>
<td>497</td>
</tr>
<tr>
<td>[Cu(3)2]</td>
<td>7</td>
<td>6.73</td>
<td>541</td>
</tr>
<tr>
<td>[Cu(4)2]</td>
<td>7</td>
<td>5.07</td>
<td>605</td>
</tr>
<tr>
<td>[Cu(5)2]</td>
<td>7</td>
<td>2.74</td>
<td>634</td>
</tr>
<tr>
<td>N719</td>
<td>14.80</td>
<td>673</td>
<td>0.61</td>
</tr>
</tbody>
</table>

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Oxidation of the copper(I) ion in [Cu(5)]⁺ occurs at lower potential than in [Cu(1)]⁺, and the value of +0.36 V is similar to that reported for [CuL₂]⁺ where L = 4,4′,6,6′-tetraphenyl-2,2′-bipyridine (+0.39 V vs. Fe/Fc⁺ in CH₂Cl₂, corrected from the published value⁵¹ of +0.77 V vs. SCE). The structural reorganization that accompanies metal oxidation is less pronounced in [Cu(5)]⁺ than in [Cu(1)]⁺. This assumption is based upon the flattened geometry in the former brought about by π-stacking interactions that occur between the phenyl groups of one ligand and the bpy domain of the other. This has been observed in a number of related structures.3⁷–⁴¹ Each of the complexes containing aliphatic substituents exhibits an irreversible oxidation process at higher potential (Table 1) most likely associated with the oxidation of the 4-bromophenyl substituent to the corresponding radical cation.⁵²

Performance of the dyes in DSCs

Heteroleptic dyes were assembled by stepwise adsorption of the anchoring ligand 6 or 7 onto TiO₂ followed by treatment with the labile⁵³,⁵⁴ homoleptic complexes [CuL₂]PF₆ (L = 1–5) resulting in ligand exchange and surface immobilisation of the heteroleptic dyes [Cu(L)(6)]⁺ and [Cu(L)(7)]⁺ shown in Scheme 2. The proximity to the surface of the copper(I) centre and of the 6,6'-substituents in the ancillary ligand is related to the presence or not of the aryl spacer in the anchoring ligand. All cells were fully masked to prevent overestimation of conversion efficiency.⁵⁵ Table 2 summarizes the DSC characteristics measured for each cell over a period of 4 days and with respect to the standard dye N719 (in the table, Voc = open circuit voltage, J₈C = short circuit current density, ff = fill factor). The final column of the table shows the energy conversion efficiencies relative to N719 which is arbitrarily set to 100%

The surface-bound heteroleptic dyes [Cu(1)(6)]⁺, [Cu(2)(6)]⁺, [Cu(3)(6)]⁺ and [Cu(4)(6)]⁺ show a gain in values of efficiency, reaching 25.2% with respect

### Table 3

DSC performance data for an independent set of sealed and masked cells using anchoring ligand 7, data are with respect to standard dye N719 measured under the same conditions.

<table>
<thead>
<tr>
<th>[CuL₂]⁺ Anchor</th>
<th>JSC/mA cm⁻²</th>
<th>Voc/mV</th>
<th>ff</th>
<th>η/%</th>
<th>Relative η/%</th>
</tr>
</thead>
<tbody>
<tr>
<td>[Cu(1)]⁺ 7</td>
<td>4.77</td>
<td>503</td>
<td>0.73</td>
<td>1.76</td>
<td>24.9</td>
</tr>
<tr>
<td>[Cu(2)]⁺ 7</td>
<td>4.22</td>
<td>482</td>
<td>0.74</td>
<td>1.49</td>
<td>21.1</td>
</tr>
<tr>
<td>[Cu(3)]⁺ 7</td>
<td>6.42</td>
<td>555</td>
<td>0.73</td>
<td>2.60</td>
<td>36.7</td>
</tr>
<tr>
<td>[Cu(4)]⁺ 7</td>
<td>4.80</td>
<td>502</td>
<td>0.72</td>
<td>1.74</td>
<td>24.6</td>
</tr>
<tr>
<td>[Cu(5)]⁺ 7</td>
<td>3.97</td>
<td>605</td>
<td>0.69</td>
<td>1.67</td>
<td>23.6</td>
</tr>
<tr>
<td>N719</td>
<td>15.37</td>
<td>648</td>
<td>0.71</td>
<td>7.08</td>
<td>100.0</td>
</tr>
</tbody>
</table>

### Table 4

On the day of sealing the cell

<table>
<thead>
<tr>
<th>[CuL₂]⁺ Anchor</th>
<th>JSC/mA cm⁻²</th>
<th>Voc/mV</th>
<th>ff</th>
<th>η/%</th>
<th>Relative η/%</th>
</tr>
</thead>
<tbody>
<tr>
<td>[Cu(1)]⁺ 7</td>
<td>6.09</td>
<td>520</td>
<td>0.74</td>
<td>2.33</td>
<td>28.4</td>
</tr>
<tr>
<td>[Cu(2)]⁺ 7</td>
<td>5.61</td>
<td>499</td>
<td>0.74</td>
<td>2.06</td>
<td>25.1</td>
</tr>
<tr>
<td>[Cu(3)]⁺ 7</td>
<td>6.74</td>
<td>571</td>
<td>0.73</td>
<td>2.81</td>
<td>34.2</td>
</tr>
<tr>
<td>[Cu(4)]⁺ 7</td>
<td>5.92</td>
<td>510</td>
<td>0.73</td>
<td>2.19</td>
<td>26.6</td>
</tr>
<tr>
<td>[Cu(5)]⁺ 7</td>
<td>6.81</td>
<td>611</td>
<td>0.72</td>
<td>3.01</td>
<td>36.6</td>
</tr>
<tr>
<td>N719</td>
<td>17.10</td>
<td>677</td>
<td>0.71</td>
<td>8.22</td>
<td>100.0</td>
</tr>
</tbody>
</table>

### Table 5

1 day after sealing the cell

<table>
<thead>
<tr>
<th>[CuL₂]⁺ Anchor</th>
<th>JSC/mA cm⁻²</th>
<th>Voc/mV</th>
<th>ff</th>
<th>η/%</th>
<th>Relative η/%</th>
</tr>
</thead>
<tbody>
<tr>
<td>[Cu(1)]⁺ 7</td>
<td>5.93</td>
<td>525</td>
<td>0.74</td>
<td>2.31</td>
<td>27.8</td>
</tr>
<tr>
<td>[Cu(2)]⁺ 7</td>
<td>5.42</td>
<td>501</td>
<td>0.73</td>
<td>1.99</td>
<td>24.0</td>
</tr>
<tr>
<td>[Cu(3)]⁺ 7</td>
<td>6.91</td>
<td>576</td>
<td>0.72</td>
<td>2.87</td>
<td>34.6</td>
</tr>
<tr>
<td>[Cu(4)]⁺ 7</td>
<td>5.51</td>
<td>521</td>
<td>0.73</td>
<td>2.09</td>
<td>25.2</td>
</tr>
<tr>
<td>[Cu(5)]⁺ 7</td>
<td>6.96</td>
<td>606</td>
<td>0.74</td>
<td>3.10</td>
<td>37.4</td>
</tr>
<tr>
<td>N719</td>
<td>17.2</td>
<td>684</td>
<td>0.71</td>
<td>8.30</td>
<td>100.0</td>
</tr>
</tbody>
</table>

### Table 6

2 days after sealing the cell

<table>
<thead>
<tr>
<th>[CuL₂]⁺ Anchor</th>
<th>JSC/mA cm⁻²</th>
<th>Voc/mV</th>
<th>ff</th>
<th>η/%</th>
<th>Relative η/%</th>
</tr>
</thead>
<tbody>
<tr>
<td>[Cu(1)]⁺ 7</td>
<td>5.73</td>
<td>530</td>
<td>0.74</td>
<td>2.26</td>
<td>26.8</td>
</tr>
<tr>
<td>[Cu(2)]⁺ 7</td>
<td>5.51</td>
<td>513</td>
<td>0.74</td>
<td>2.09</td>
<td>24.7</td>
</tr>
<tr>
<td>[Cu(3)]⁺ 7</td>
<td>6.69</td>
<td>584</td>
<td>0.73</td>
<td>2.85</td>
<td>33.7</td>
</tr>
<tr>
<td>[Cu(4)]⁺ 7</td>
<td>5.44</td>
<td>531</td>
<td>0.71</td>
<td>2.05</td>
<td>24.3</td>
</tr>
<tr>
<td>[Cu(5)]⁺ 7</td>
<td>7.03</td>
<td>590</td>
<td>0.73</td>
<td>3.01</td>
<td>35.6</td>
</tr>
<tr>
<td>N719</td>
<td>17.23</td>
<td>696</td>
<td>0.71</td>
<td>8.45</td>
<td>100.0</td>
</tr>
</tbody>
</table>

### Table 7

7 days after sealing the cell

<table>
<thead>
<tr>
<th>[CuL₂]⁺ Anchor</th>
<th>JSC/mA cm⁻²</th>
<th>Voc/mV</th>
<th>ff</th>
<th>η/%</th>
<th>Relative η/%</th>
</tr>
</thead>
<tbody>
<tr>
<td>[Cu(1)]⁺ 7</td>
<td>5.73</td>
<td>530</td>
<td>0.74</td>
<td>2.26</td>
<td>26.8</td>
</tr>
<tr>
<td>[Cu(2)]⁺ 7</td>
<td>5.51</td>
<td>513</td>
<td>0.74</td>
<td>2.09</td>
<td>24.7</td>
</tr>
<tr>
<td>[Cu(3)]⁺ 7</td>
<td>6.69</td>
<td>584</td>
<td>0.73</td>
<td>2.85</td>
<td>33.7</td>
</tr>
<tr>
<td>[Cu(4)]⁺ 7</td>
<td>5.44</td>
<td>531</td>
<td>0.71</td>
<td>2.05</td>
<td>24.3</td>
</tr>
<tr>
<td>[Cu(5)]⁺ 7</td>
<td>7.03</td>
<td>590</td>
<td>0.73</td>
<td>3.01</td>
<td>35.6</td>
</tr>
<tr>
<td>N719</td>
<td>17.23</td>
<td>696</td>
<td>0.71</td>
<td>8.45</td>
<td>100.0</td>
</tr>
</tbody>
</table>
Density, for both \([\text{Cu}(\text{dpp})_2\text{(dtbbpy})]\) and \([\text{Cu}(\text{dpp})_2\text{Ph}]\), is an additional degree of freedom associated with the rotation about the interannular C–C bond of the bpy ligand.

In order to check the reproducibility of our results, a new set of masked solar cells was prepared using anchoring ligand 7, and cell performances were measured over a 7 day period (Table 3). A comparison of the data in Tables 2 and 3 reveals analogous trends, although absolute efficiencies are lower for the second batch of cells than for the first. Fig. 10a shows plots of current density against open-circuit voltage for \([\text{Cu}(3)(7)]\), illustrating the initial maturing of the cell, followed by consistent behaviour for a week. The ripening process of the cell containing \([\text{Cu}(5)(7)]\) is illustrated in the I–V curve in Fig. 10b and replicates that observed for the first set of cells. Good fill factors were obtained for all cells in both sets (Tables 2 and 3).

The incident photon-to-current conversion efficiency (IPCE) or external quantum efficiency (EQE) spectra were measured for each of the solar cells detailed in Table 3. The spectra shown in Fig. 11a for \([\text{Cu}(3)(7)]\) are representative of those of cells containing the dyes \([\text{Cu}(1)(7)]\), \([\text{Cu}(2)(7)]\), \([\text{Cu}(3)(7)]\) and \([\text{Cu}(4)(7)]\). The peak maxima in the IPCE spectra at \(\approx 470\) nm correlate with the observed maxima (\(\lambda_{\text{max}} = 481–488\) nm) in the solution absorption spectra of the homoleptic complexes \([\text{Cu}(1)_2]\), \([\text{Cu}(2)_2]\), \([\text{Cu}(3)_2]\) and \([\text{Cu}(4)_2]\) (Fig. 9). IPCE peak maxima of between 39 and 42% are achieved for \([\text{Cu}(1)(7)]\), 38–40% for \([\text{Cu}(2)(7)]\), 47–50% for \([\text{Cu}(3)(7)]\) and 38–39% for \([\text{Cu}(4)(7)]\), corresponding to the trends in efficiencies given in Table 3. Fig. 11b shows the IPCE spectra for \([\text{Cu}(5)(7)]\) from the day of sealing to a week after sealing the cell. Initially an IPCE maximum of 28% is observed but, consistent with the ripening of the cell described above, this increases to 46% after a week. It is noticeable that the shape of the IPCE curve for \([\text{Cu}(5)(7)]\) is close to those of the other complexes, i.e. the differences observed in the solution absorption spectra between the homoleptic complexes containing alkyl and phenyl substituents are not replicated in the IPCE curves of the heteroleptic complexes. To gain more insight into this, the absorption spectra of the surface-anchored heteroleptic dyes \([\text{Cu}(3)(7)]\) and \([\text{Cu}(5)(7)]\) were recorded. The anodes were prepared as described in the Experimental section (solar cell fabrication), and their absorption spectra are shown in Fig. 12. The absorption spectrum of the photoanode with adsorbed \([\text{Cu}(5)(7)]\) also exhibits an enhanced spectral response at ca. 600 nm as also seen in the homoleptic complex \([\text{Cu}(5)_2]\) in solution (Fig. 9).

Fig. 13 compares the IPCE spectra of cells containing the five copper(i)-containing dyes with that of N719. The IPCE curves indicate improved response in the red region of the spectrum consistent with the red-shifting of the absorption spectra of the dyes containing ancillary ligand 5. While good photon-to-current quantum efficiency is achieved with the copper(i) dyes over the range 450–550 nm, the dyes do not
achieve the broad spectra response exhibited by the state-of-the-art ruthenium(II) dye N719. Specifically, aspects of molecular design that need to be addressed include the further red-

shifting of the absorption and also an increase in the absorptivity.

**Conclusions**

We have described the syntheses and characterizations of the homoleptic complexes [CuL₂][PF₆] with L = 1–5. The solid state structures of 2[[Cu(1)₂][PF₆]]·3Me₂CO, [Cu(2)₂][PF₆], and 2[[Cu(3)₂][PF₆]]·Et₂O confirm distorted tetrahedral environments for the Cu⁺ ions. In [Cu(5)₂][PF₆]·CH₂Cl₂, the cation possesses a significantly flattened geometry caused by π-interactions between the pendant phenyl rings and bpy domains of adjacent ligands. The solution absorption spectrum of each of [Cu(1)₂][PF₆], [Cu(2)₂][PF₆], [Cu(3)₂][PF₆] and [Cu(4)₂][PF₆] shows an MLCT band at around 485 nm. Greater absorptivity in the red region of the spectrum is achieved by replacing the 6- and 6'-alkyl substituents with phenyl groups in the complex [Cu(5)₂][PF₆] and results in the appearance of two broad bands in the visible region, the longer wavelength band at 573 nm.

The efficiencies, I–V curves and IPCE spectra of DSCs containing the anchored dyes (a) [Cu(3)⁺(7)⁺] and (b) [Cu(5)⁺(7)⁺] over a period of a week after sealing the solar cells.

**Acknowledgements**

We thank the European Research Council (Advanced Grant 267816 LiLo), the Swiss National Science Foundation, the
Swiss Nanoscience Institute and the University of Basel for financial support. Highfield NMR spectra were recorded by Nik Hostettler, Cathrin Ertl or Dr. Colin Martin; Ewald Schönhofer and Annika Bütten prepared and screen-printed the mesoporous TiO₂. Liselotte Siegfried is thanked for assisting with UV-vis measurements of the surface-anchored dyes.

Notes and references